BALABHADRA SKILL DEVELOPMENT ACADEMY
SCIENCE QUESTION BANK - 14

Time: 1 Hour   Full marks: 59   Pass marks: 47

1.   The _____ of a solid is an indication of the strength of the force of attraction between its particles.
2.   The melting point of ice is ____ .
3.   The kinetic energy of particles is minimum in_____, intermediate in ____and the maximum in _____.
4.   In case of _____ , the arrangement of particles is most ordered.
5.   In case of _____ , layers of particles can slip and slide over each other.
6.   _____ is a surface phenomenon.
7.   The rate of evaporation increases with increase in _____, _____ and _____.
8.   The rate of evaporation decreases with decrease in _____.
9.   The amount of heat required to convert unit mass of the solid into the liquid without a change in temperature is called ____ .
10. The amount of heat required to convert unit mass of a liquid into the vapour state at its boiling point without a change in temperature is called _____ .
11. The intermixing of particles of two different types of matter on their own is called _____ .
11(a) Diffusion becomes faster with increase in _____.
12. The temperature at which a liquid starts boiling at the atmospheric pressure is known as _____ .
13. A pure solid will have a ____ melting point.
14. The temperature of a pure liquid will remain ____ as its boiling point.
15. A pure substance will produce only one well-defined spot on _____.
16. The amount of solute present in grams in one litre of the solution is called _____.
17. In _____ only physical properties of the substances change but no new substances are formed.
18. In _____ in which new substances are formed and chemical properties of substances changes.
18(a) Rusting of iron is _____ change.
19. The phenomenon of scattering of light by the particles (colloid or suspension) due to which the path of light becomes visible is called ____ .
20. The maximum amount of a solute which can be dissolved in 100g of a solvent at a specified temperature is known as _____.
21. A solution in which no more solute can be added in the particular amount of solvent is called _____.
22. All matter is made of very tiny particles called _____ .
23. _____ are indivisible particles, which cannot be created or destroyed in a chemical reaction.
24. Atoms of a given element are identical in ____ and ____ properties.
25. Atoms of different elements have different _____ and _____ properties.
26. Atoms combine in the ratio of _____ to form compounds.

27. The relative number and kinds of atoms are _____ in a given compound.
28. The energies of permissible orbits of atoms are fixed and are called _____.
29. Energy is emitted or absorbed by an atom only when an _____ moves from one orbit another.
30. Atoms of the same element having same atomic number but different mass number is called _____.
31. The atoms of different elements having different atomic numbers but same mass number is called _____.
32. Cell was first discovered by _____ in 1665.
33. _____ and _____ are semi-autonomous organelles as both of them have their genome (DNA) and ribosomes to synthesis their own proteins.
34. _____ are functional units of DNA.
34(a) _____ carry parental hereditary information to next generation.
35. _____ occur during cell division and they are formed due to _____ and _____ of chromes.
36. Most plant Cells have a large _____ that helps to maintain the _____ of the cell stores important substances including wastes.
37. _____ is a thin layer of cementing substance.
37(a) Middle lamela contain _____, _____ and _____.
38. The seeds of cotton contain numerous _____ which form the husk of cotton.
39. Cells of cork are _____ and compactly arranged without _____.
40. _____ are tubular cells with perforated walls.
41. The cell generally associated with sieve tubes is a small thin-walled cell containing dense and active cytoplasm with a large elongated nucleus is called _____ cell.
42. _____ anchors the muscles and supports the main organs of the body.
43. Bone Cells are embedded in a hard matrix that is composed of ___ and _____ compound.
44. In the respiratory tract, the columnar epithelial tissue has___, which are hair like projection on outer surfaces of epithelial cells.
45. Classification helps us in exploring the _____ of life forms.
46. Organisms are made of ___ or ___ cells.
47. The cells have a _____.
48. The _____ nomenclature makes for a uniform way of identification of the vast diversity around us.
49. _____ classified organisms depending on their habitat.
50. The force required to stop a moving body is directly proportional to its _____ and _____.
51. The state when an object does not weigh anything is called _____.
52. When both the objects are very big having very large masses, then _____ between them becomes extremely large.
53. Energy shells of atoms are designated as _____.
54. First energy shell is designated by _____ and n=_____.
55. Second energy shell is designated by _____ and n= _____.

1.   Melting point
2.   273.16 K
3.   Solid, liquid, gas
4.   Solids
5.   Liquids
6.   Evaporation
7.   Surface area, temperature, wind speed
8.   Humidity
9.   Latent Heat of Fusion
10.   Latent Heat of Vaporisation
11.   Diffusion
11.(a) Temperature
12.   boiling point
13.   sharp
14.   steady
15.   chromatogram
16.   Strength of solution
17.   Physical change
18.   Chemical change
18.(a) Chemical
19.   Tyndall effect
20.   Solubility, solute
21.   Saturated Solution
22.   atoms
23.   Atoms
24.   mass, chemical
25.   mass, chemical
26.   small whole numbers
27.   Constant
28.   Station states or energy levels or shells
29.   electron
30.   Isotopes
31.   Isobars
32.   Robert Hooke
33.   Mitochondria, Chloroplasts
34.   Genes
34(a)   Genes
35.   Chromosomes, coiling, folding
36.   central vacuole, ~~turgity~~ turdigity
37.   Middle Lamella
37(a)   Pectin, Lignin, Protein
38.   long unicellular hair
39.   dead, intercellular space
40.   Sieve tubes

41. companion
42. Bone
43. calcium, phosphorous
44. cilia
45. diversity
46. prokaryotic, eukaryotic
47. cell-wall
48. binomial
49. Aristotle
50. mass, velocity
51. weightlessness
52. gravitational force of attraction
53. K,L,M,N,----
54. K, n=1
55. L, n=2