BALABHADRA SKILL DEVELOPMENT ACADEMY
SCIENCE QUESTION BANK - 8

Time: 1 Hour Full marks: 56 Pass marks: 44

1. Mixture can be separated into _____ using appropriate separation techniques.
2. A solution is a _____ mixture of two or more substances.
3. The major component of a solution is called the_____, and minor the _____.
4. The concentration of a solution is the amount of solute present per unit _____ or per unit _____ of the solution.
5. Material that are insoluble in a solvent and have particles that are visible to naked eyes, form a _____ .
6. A suspension is a _____ mixture.
7. _____ are heterogeneous mixtures in which the particle size is too small to be seen with the naked eye, but is big enough to _____ light.
8. Colloids are useful in _____ and _____ .
9. In colloids the particles are called _____ and the medium in which they are distributed is called the _____.
10. Pure substances can be _____ or_____ .
11. A Compound is a substance composed of _____ and chemically combined in a _____.
12. Properties of a _____ are different from its constituent elements, whereas a _____ shows the properties of its constituting elements or compounds.
13. Pure Substance has the _____, _____ , _____ and _____ at a give temperature and pressure.
14. A pure substance can be either an _____ or a _____.
15. During a Chemical reaction, _____ remains unchanged and this is known as _____.
16. In a pure chemical compound, elements are always present in a _____ and this is known as the Law of_____.
17. An _____ is the smallest particle of the element

17(a) _____ cannot usually exist independently.

17(b) Atom retains all _____ of element.

18. A _____ is the smallest particle of an element or a compound.

18(a) _____ is capable of independent existence under ordinary conditions.

19. _____ shows all the properties of the substance.
20. A Chemical formula of a compound shows _____ and _____.
21. Cluster of atoms that act as an ion are called _____ .

21(a) Polyatomic ions carry _____ on them.

22. The chemical formula of a molecular compound is determined by the _____ of each element.
23. In ionic compounds, the _____ on each ion is used to determine the chemical formula of the compound.
24. Scientists use the _____to compare the masses of different atoms of elements.
25. Atoms of Carbon-12 isotopes are assigned a relative atomic mass of _____ .

26. The Avogadro constant is _____ .

26(a) The Avogadro constant is defined as the number of atoms in exactly _____.

27. The _____ is the amount of substance that contains the same number of particles (atoms/ ions/ molecules/ formula units etc.) as there are atoms in exactly 12 g of carbon-12.

27(a) Mole of a substance contains _____ no. of atoms, molecules or ions.

28. Mass of one mole of a substance expressed in gms is called its _____ .

29. _____ is a form of matter that cannot be broken down by chemical reactions into simpler substances.

30. A mixture is formed when two or more substances are mixed in _____.

31. If the mixture has a uniform composition thorough out than it is called a _____ mixture.

32. If the mixture has physically distinct part and has non-uniform compositions, then it is called _____ mixture.

33. Sugar in water is a _____.

34. Alloy of Copper and Zinc is a _____.

35. Ethyl alcohol in water is a _____.

36. When solids are dispersed in liquids to form a Heterogeneous mixture or an opaque medium, it is called _____.

37. Mass percentage of solution is _____.

38. Volume percentage of solution is _____.

39. A true solution is a _____, colloidal solution _____ and suspension is _____.

40. Size of solute in true solution is _____.

41. Size of solute in colloidal solution is _____.

42. Size of solute in suspension is _____.

43. According to _____ theory, elements or compounds are made up of small particles, atoms or molecule.

44. _____ is the smallest particle of an element.

45. Atom retains _____ through out all physical and chemical change.

46. _____ is a group of two or more atoms chemically bonded together.

47. _____ is smallest part of element or compound.

48. _____ is a charged particle.

49. The combining capacity of an element is known as _____.

50. The relative masses of atoms are given in _____.

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SCIENCE QUESTION BANK – 8 (ANSWER)

1. pure substances
2. homogeneous
3. solvent, solute
4. volume, mass
5. suspension
6. heterogeneous
7. Colloids, scatter
8. industry, daily life
9. dispersed phase, dispersion medium
10. elements or compounds
11. two or more different types of elements, fixed proportion
12. compound, mixture
13. same colour, composition, taste, texture
14. element or compound
15. sum of masses of reactants and products, law of conversation of masses
16. Definite Proportion by mass, Definite Proportions
17. atom

17(a) atom chemical properties

18. molecule

18(a) molecule

19. molecule
20. constituent elements, number of atoms of each combining element
21. polyatomic ions

21(a) fixed charge

22. valency
23. charge
24. relative atomic mass scale
25. 12

25(a) a carbon-12 atom

26. 6.022×10^{23}

26(a) 12g of carbon-12 atom

27. mole

27(a) 6.022×10^{23}

28. molar mass
29. Element
30. any proportion without any chemical change
31. Homogeneous
32. Heterogeneous
33. Homogeneous mixture
34. Homogeneous mixture
35. Homogeneous mixture
36. Suspension
37. (Mass of solute /Mass of Solution) X 100
38. (Mass of solute /Volume of Solution) X 100

39. Homogeneous mixture, appears to be Homogeneous, but Heterogeneous mixture, Heterogeneous mixture
40. less than 10^{-8} cm
41. between 10^{-7} to 10^{-5} cm
42. Larger than 10^{-5} cm
43. Daltons atomic theory
44. Atom
45. Identity
46. Molecule
47. Molecule
48. Ion
49. Valency
50. atomic mass unit